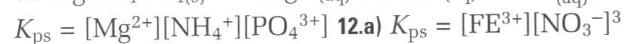
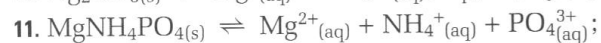
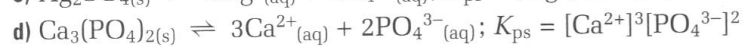
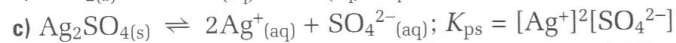


## Chimie 40S

### Devoir : la solubilité – solutions

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b) plus petit 13.  $2,2 \times 10^{-16}$  14.  $1,1 \times 10^{-37}$  15.  $3,2 \times 10^{-11}$

16.a)  $4,8 \times 10^{-29}$  b) 1 g de solution = 1 mL de solution

17.a)  $1,3 \times 10^{-5}$  mol/L b)  $7,8 \times 10^{18}$  unités de formule de

AgCl/L c)  $1,9 \times 10^{-4}\%$  (m/v) 18.  $1,0 \times 10^{-10}$  mol/L

19.  $9,9 \times 10^{-3}$  mol/L; 4,1 g/L

20.  $8,5 \times 10^{12}$  unités de formule de ZnS/L

21.a)  $1,33 \times 10^{-5}$  mol/L b)  $1,2 \times 10^{-9}$  mol/L

22.  $3,9 \times 10^{-6}$  mol/L 23.a)  $7,018 \times 10^{-3}$  mol/L

b)  $1,97 \times 10^{-4}$  mol/L 24.a)  $1,6 \times 10^{-2}$  mol/L

b)  $4,0 \times 10^{-4}$  mol/L :