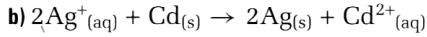
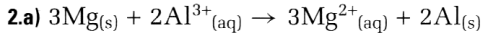
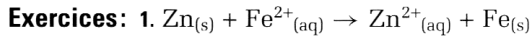


Chimie 40S

Devoir : oxydoréduction – solutions

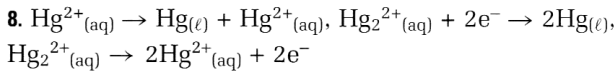
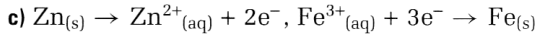
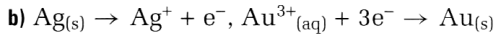
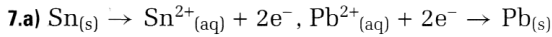
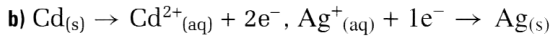
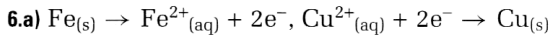
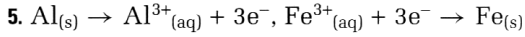


3.a) Mg oxydé, Al^{3+} réduit

b) Cd oxydé, Ag⁺ réduit

4.a) Al^{3+} agent oxydant, Mg agent réducteur

b) Ag⁺ agent oxydant, Cd agent réducteur



9.a) +3 b) 0 c) +6 d) +5 e) 0 f) +2

10.a) H, +1; S, +4; O, -2

b) H, +1; O, -2 c) H, +1; P, +5; O, -2

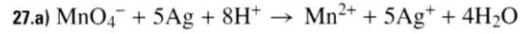
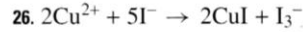
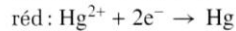
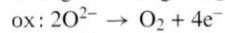
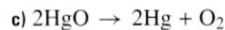
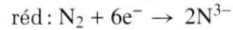
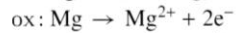
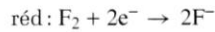
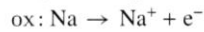
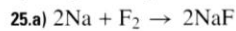
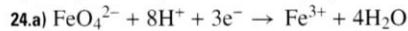
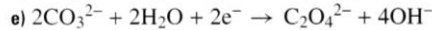
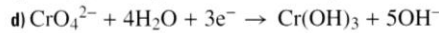
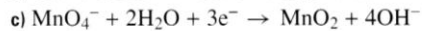
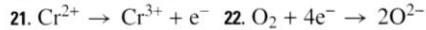
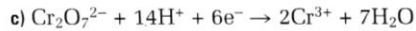
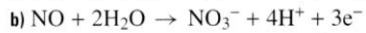
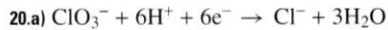
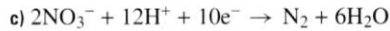
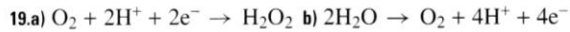
11.a) +2 b) -1 12.a) Al, +3; H, +1; C, +4; O, -2

b) N, -3; H, +1; P, +5; O, -2 c) K, +1; H, +1; I, +7; O, -2

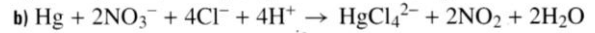
13.a) oui b) non

14.a) H_2O_2 agent oxydant, Fe^{2+} agent réducteur

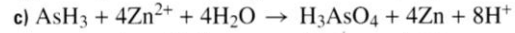
15. ClO_2^{-} oxydé, Br_2 réduit 16. oui



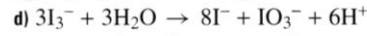
agent oxydant, MnO_4^{-} ; agent réducteur, Ag



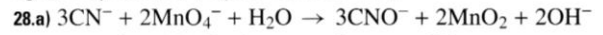
agent oxydant, NO_3^{-} ; agent réducteur, Hg



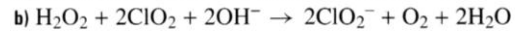
agent oxydant, Zn^{2+} ; agent réducteur, AsH_3



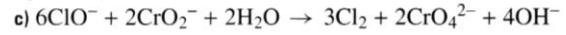
agent oxydant, I_3^{-} ; agent réducteur, I_3^{-}



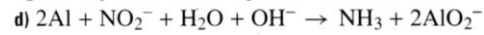
agent oxydant, MnO_4^{-} ; agent réducteur, CN^{-}



agent oxydant, ClO_2 ; agent réducteur, H_2O_2



agent oxydant, ClO^{-} ; agent réducteur, CrO_2^{-}



agent oxydant, NO_2^{-} ; agent réducteur, Al

