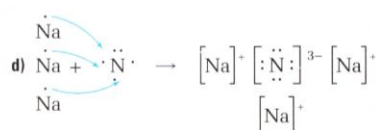
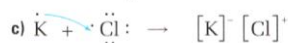
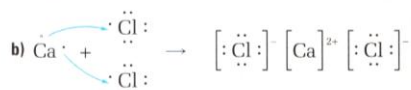
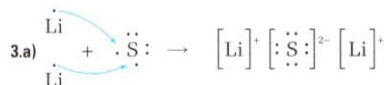


Chimie 40S

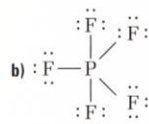
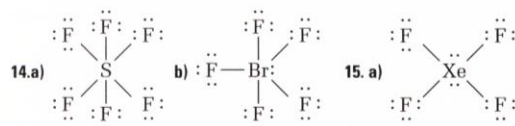
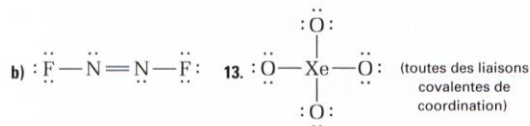
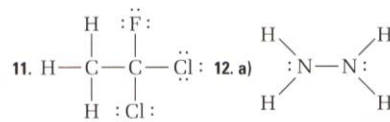
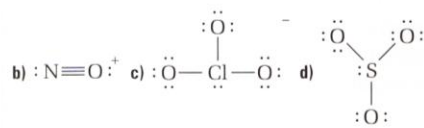
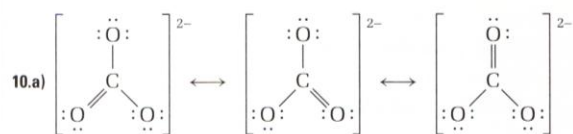
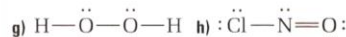
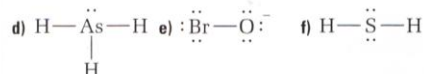
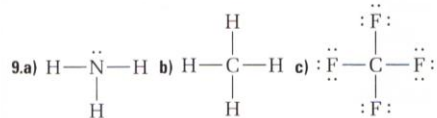
Devoir : liaisons – solutions

1. a) $1s^2$ b) $1s^2 2s^2 2p^6 3s^2 3p^6$ c) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6$

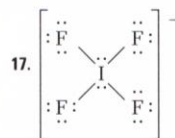
d) $1s^2 2s^2 2p^6$ 2.a) $[\text{Li}]^+$ b) $[\text{Ca}]^{2+}$ c) $[\text{Br}]^-$ d) $[\text{O}]^{2-}$



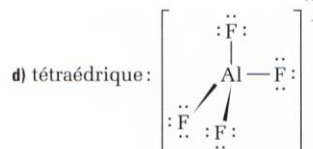
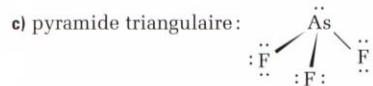
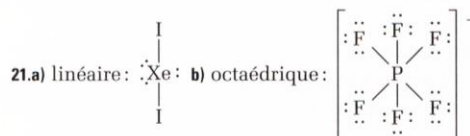
4.a) 6 (VIA) b) 1 (IA) c) 3 (IIIA) d) 2 (IIA) 5. H—Cl, H—Br, H—I 6. AsBr_3 $\Delta EN = 0,78$, CaBr_2 $\Delta EN = 1,96$, KBr $\Delta EN = 2,14$: une augmentation du caractère ionique 7.a) C—H b) Sn—F c) C—O 8. covalente: S_8 , F_2 ; covalente polaire $\text{SCl}_2 < \text{SF}_2 < \text{PF}_3$; ionique: RbCl



16. Dans PI_3 , un doublet libre autour de P; dans ClI_3 , deux doublets libres autour de Cl.



18.a) linéaire b) angulaire c) triangulaire plane d) tétraédrique
19.a) tétraédrique b) bipyramide triangulaire c) tétraédrique d) tétraédrique 20. NO_2^+ linéaire; NO_2^- courbée



22.a) O b) P c) Cl 23.a) tétraédrique, polaire b) triangulaire plane, polaire c) triangulaire plane, non polaire 24. tétraédrique, polaire 25.a) NF_3 b) TeCl_4 26. pyramide triangulaire, en forme de T